

Chapter 6 Quiz

Part A: Modified True/False

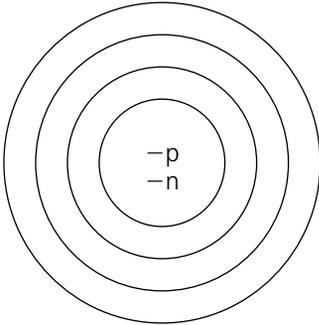
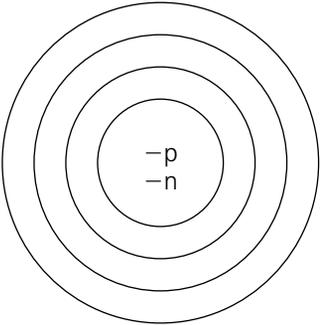
Indicate in the left-hand column whether each statement is true or false. If the statement is false, change the statement to make it true.

____ 1. A compound forms when two or more substances are put together, but are not chemically combined. _____

____ 2. An electron is a negatively charged subatomic particle. _____

Part B: Diagrams

In the spaces below, draw Bohr diagrams for the elements indicated.

<p>3. oxygen</p>  <p>element _____ atomic number _____</p>	<p>4. magnesium</p>  <p>element _____ atomic number _____</p>
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Part C: Multiple Choice

Circle the letter beside the answer that best completes the statement or answers the question.

5. Which of the following elements has the smallest atomic mass?

- A. helium
- B. oxygen

- C. nitrogen
- D. hydrogen

6. If an atom has 10 protons, how many electrons does it have in its second electron shell?

- A. 2
- B. 8

- C. 10
- D. 20

Chapter 6 Quiz (continued)

7. What is milk an example of?
- A. mixture
B. molecule
C. compound
D. pure substance
8. A Bohr diagram shows an element with 28 protons. Which element does the Bohr diagram represent?
- A. nickel
B. silicon
C. barium
D. nitrogen
9. Which of the following is an example of a compound?
- A. F_2
B. Ne
C. Cu^{2+}
D. NaCl
10. How many electrons does an element in Group 14 on the Periodic Table have in its outer shell?
- A. 2
B. 4
C. 8
D. 10
11. An element that has 3 electron shells will be found in the same period on the Periodic Table as which element?
- A. tin
B. sulfur
C. carbon
D. bromine
12. Which of the following is an example of a pure substance?
- A. fog
B. soil
C. gold
D. glue

Part D: Short Answer

13. Copper has an atomic number of 29 and an atomic mass of 63.546 *u*.

a. Why isn't copper's atomic mass closer to 58 *u*?

b. About how many neutrons does the average atom of copper have?
