PERIODIC TABLE WORKSHEET

Name _____

1. Where are the most active metals located? _____
2. Where are the most active nonmetals located? _____
3. As you go from left to right across a period, the atomic size (decreases / increases). Why? _____
4. As you travel down a group, the atomic size (decreases / increases). Why?

   _____
5. A negative ion is (larger / smaller) than its parent atom.
6. A positive ion is (larger / smaller) than its parent atom.
7. As you go from left to right across a period, the first ionization energy generally (decreases / increases). Why? _____
8. As you go down a group, the first ionization energy generally (decreases / increases) Why? _____
9. Where is the highest electronegativity found? _____
10. Where is the lowest electronegativity found? _____
11. Elements of Group 1 are called _____
12. Elements of Group 2 are called _____
13. Elements of Group 3-12 are called _____
14. As you go from left to right across the periodic table, the elements go from (metals / nonmetals) to (metals / nonmetals).
15. Group 17 elements are called _____
16. The most active element in Group 17 is _____
17. Group 18 elements are called _____
18. What sublevels are filling across the Transition Elements? _____
19. Elements within a group have a similar number of _____
20. Elements across a series have the same number of _____
21. A colored ion generally indicates a _____
22. As you go down a group, the elements generally become (more / less) metallic.
23. The majority of elements in the periodic table are (metals / nonmetals).
24. Elements in the periodic table are arranged according to their _____
25. An element with both metallic and nonmetallic properties is called a _____

6.3 Name: _____ Date: _____

Analyzing the Periodic Table

Answer the questions below about the elements in the Periodic Table.

1. In what Group of the Periodic Table would you most likely find a gas with atoms that have 8 electrons in their outermost shells? _____
2. How many electron shells does an element in Period 7 have? _____
3. Which group on the Periodic Table has the widest range of properties? _____
4. Complete the table below to describe elements in different groups. _____

	Alkali Metals	Alkaline Earth Metals	Halogens	Noble Gases
Number/range of electrons in outer shell				
Number/range of electron shells				

5. Using information from the table above, identify which characteristic—number of electrons in outer shell or number of electron shells—seems to have the bigger effect on the properties of elements. Explain your answer.

_____

_____

_____

_____

_____

6. Alkali metals and halogens readily form ions, but the noble gases do not. Using information from the table above, explain why.

_____

_____

_____

_____

_____